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Chapter 12 Stoichiometry

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Chemical Calculations Answer

Stoichiometry Formula Weight

(FW)! • Sum of the atomic weights for the atoms in a

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chemical formula • So, the formula weight of calcium chloride, CaCl_2 , would be $\text{Ca}: 1(40.1 \text{ amu}) + \text{Cl}: 2(35.5 \text{ amu}) = 111.1 \text{ amu}$ • These are generally Page 5/28

Stoichiometry Workbook

Answers

Calculate the number of moles of each reactant present: 5.272 mol of TiCl_4 and 8.23 mol of Mg. Divide the actual number of moles of each reactant by its stoichiometric coefficient in the balanced chemical equation:

$$\begin{aligned} \text{TiCl}_4: & \frac{5.272 \text{ mol(actual)}}{1 \text{ mol(stoich)}} = 5.272 \\ \text{Mg}: & \frac{8.23 \text{ mol(actual)}}{2 \text{ mol(stoich)}} = 4.12 \end{aligned}$$

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Chapter 03 - Stoichiometry

Stoichiometry is a field of chemistry focused on the ratio of reactants and products in a chemical reaction. Different units can be used in these ratios, such as mass (grams, kilograms, tonnes, ...

Who discovered

stoichiometry? | Study.com

When 80 grams of aluminum is reacted with excess chlorine gas, how many formula units of AlCl_3 are produced? $\times 1$
mole Al = 2.96 moles Al
There is a 1:1 ratio between Al and AlCl_3 , therefore

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there are 2.96 moles AlCl_3 .
 $= 1.78 \times 10^25$

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Answer Key. Study Guide
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Answer The study of the
quantitative relationships
between the amounts of
reactants used and the
amounts of products formed
by a chemical reaction is
called stoichiometry. _____
2. Stoichiometry is based on
the law of conservation of
mass. _____ 3.

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Stoichiometry Example C₆H₆ + Br₂ → C₆H₅Br + HBr
Benzene (C₆H₆) reacts with Bromine to produce bromobenzene (C₆H₅Br) and

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hydrobromic acid. If 30. g of benzene reacts with 65 g of bromine and produces 56.7 g of bromobenzene, what is the percent yield of the reaction? 30.g 65 g 56.7 g
...

Chapter 3 Stoichiometry - Chemistry

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Multiple Choice Questions

(MCQ) and Answers on

Stoichiometry Question 1 :

The weight fraction of methanol in an aqueous solution is 0.64. The mole fraction of methanol

X_M satisfies $X_M < 0.5$ $X_M =$

0.5 $0.5 < X_M < 0.64$ $X_M ?$

0.64 Answer : 4 Question 2 :

On addition of 1 c.c. of dilute hydrochloric acid (1% concentration) to 80 c.c. of a buffer solution of $\text{pH} = 4$, the pH of the solution becomes 1.8 ...

Stoichiometry Questions and Answers - QforQuestions

Stoichiometry is based on the law of conservation of mass, In any chemical

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reaction, the mass of the products is equal to the mass of the reactants. ... Study Guide Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.

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