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CH 7 Section 1 Chemical Names & Formulas Objectives

Explain the significance of a chemical formula Determine the formula of an ionic compound formed between two given ions Name an ionic compound given its formula Using prefixes, name a binary molecular compound from its formula Write the formula ...

CH 7 Section 1 - Google Slides

The simplest ratio of atoms of each element in a compound is called the empirical formula. Mass is conserved in chemical reactions. Part of. Chemistry (Single Science) Quantitative chemistry.

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Section Review Answers Chemical Names And Formulas

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Stoichiometry is built on an understanding of atomic masses (Section 2.4), chemical formulas, and the law of conservation of mass. (Section 2.1) The French nobleman and scientist Antoine Lavoisier (FIGURE 3.1) discovered this important chemical law during the late 1700s. Lavoisier stated the law in this eloquent way: “ We may lay it down as an ...

CHEMICAL EQUATIONS - STOICHIOMETRY: CALCULATIONS WITH ...

Chemical Formulas and Chemical Compounds SECTION 2
SHORT ANSWER Answer the following questions in the space provided. 1. Assign the oxidation number to the specified element in each of the following examples: 4 a. S in H_2SO_3 6 b. S in MgSO_4 2 c. S in K_2S 1 d. Cu in Cu_2S 6 e. Cr

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in Na_2CrO_4 5 f. N in HNO_3

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Chemical Formulas and Chemical Compounds SECTION 1
SHORT ANSWER Answer the following questions in the
space provided. 1. c In a Stock system name such as iron(III)
sulfate, the Roman numeral tells us (a) how many atoms of
Fe are in one formula unit. (b) how many sulfate ions can be
attached to the iron atom. (c) the charge on each Fe ion.

7 Chemical Formulas and Chemical Compounds
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Chemical Formulas And Compounds Chapter CHAPTER 7
REVIEW Chemical Formulas and Chemical Compounds
SECTION 1 SHORT ANSWER Answer the following questions
in the space provided. 1. c In a Stock system name such as
iron(III) sulfate, the Roman numeral tells us (a) how many
atoms of Fe are in one formula unit. (b) how many sulfate
ions

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Chemical Formulas And Compounds Chapter 7 Review
Section 1

Chapter 9 Practice Problems: Chemical Names and Formulas.

Section 9.1: Naming Ions. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e.

nickel, two electrons lost . b. iodine d. aluminum f.

magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P3

...

Chapter 9 Practice Problems: Chemical Names and Formulas

Step 1: Write the prefix of the first element (if the first

element is single, don't put a prefix) Step 2: Write the name

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of the first element Step 3: Write the prefix of the second element Step 4: Write the root of the second element Step 5: Add -ide Ex) PCl_3 Phosphorus trichloride (Tri-Clor-Ide) Ex) S_4F_6

Chapter 7 Section 7.1-7.2 (Chemical Names and Formulas ... Science, Chemical Reactions, Section 1 Chemical Formulas and Equations. STUDY. PLAY. Matter can change physically or chemically; a process that produce a chemical change is a chemical _____ reaction. A chemical _____ is a shorthand form of what reactants are used and what products are formed in a chemical reaction.

Science, Chemical Reactions, Section 1 Chemical Formulas ...

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chemical formula is a shorthand way to use chemical symbols and numbers to represent a substance. A chemical formula shows how many atoms of each kind of element are present in a molecule. The chemical formula for water, shown in Figure 1, is H_2O . This formula tells you that one water molecule is made of two hydrogen atoms and one oxygen atom.

Section 2 Chemical Formulas and Equations

Chapter 7 Section 1 Chemical Names and Formulas

Significance of a Chemical Formula, continued •

The chemical formula for an ionic compound represents one formula unit—the simplest ratio of the compound's positive ions (cations) and its negative ions (anions).

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—example: aluminum sulfate — $\text{Al}_2(\text{SO}_4)_3$...

Chapter 7 Section 2 Chemical Formulas And Compounds

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