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Draw the Lewis
structure for each
molecule. 1. PH₃

$$\begin{array}{c} \text{H} \\ | \\ \text{H} - \text{P} - \text{H} \\ | \\ \text{H} \end{array}$$

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respectively, for
single, double, and

triple P — — 2. H 2 S

H H H — H S S — 3.

HCl

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Covalent bonds
involve sharing

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electrons between atoms. When the atoms in the bond pull equally, the bonding electrons are shared equally, and the bond is nonpolar. When the atoms in the bond pull unequally, the bonding electrons are pulled closer to one atom, and the bond is polar.

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overlaps with hydrogen's 1s orbital to form a covalent bond, called σ bond.

The angle between each lobe assumes 109.47° since the sp^3 carbon is tetrahedral. sp^2

Hybrid Orbitals : If a carbon atom is in a double-bond environment, such as C_2H_4 (ethylene),

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the carbon is hybridized to become sp^2 hybrid orbital.

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Lattice energies of alkali metal iodides.

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respectively, yet their
lattice energies are
1030 and 3795
kJ/mol.

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electron sharing
usually occurs so that
atoms attain the
electron

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example, a single
hydrogen atom has
one electron. But a

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pair of hydrogen atoms shares electrons to form a covalent bond in a diatomic hydrogen molecule.

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