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Chapter 2: The Chemistry of Life

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Flashcards | Quizlet

Chapter 2: Introduction to the Chemistry of Life. Figure 2.1 Foods such as bread, fruit, and cheese are rich sources of biological macromolecules. The elements carbon, hydrogen, nitrogen, oxygen, sulfur, and phosphorus are the key building blocks of the chemicals found in living things. They form the carbohydrates, nucleic acids, proteins, and lipids (all of which will be defined later in this chapter) that are the fundamental molecular components of all organisms.

Chapter 2: Introduction to the Chemistry of Life ...

Chapter 2 The Chemistry of Life
What are the basic building blocks

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of all matter? Atoms Describe the structure of an atom protons, electrons, and neutrons. The nucleus (center) of the atom contains the protons (positively charged) and the neutrons (no charge).

Chapter 2 The Chemistry of
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Biology Chapter 2- The Chemistry of Life. Essential Question: What are the basic chemical principles that affect living things?

Biology Chapter 2- The Chemistry of Life

Biology: Chapter 2, The Chemistry of Life. Atom. Nucleus. Electron. Element. An atom is the smallest constituent unit of ordinary matter th.... The nucleus is the small,

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dense region consisting of protons
a.... The electron is a subatomic
particle, symbol e^- or $^-$, with
a....

Chapter 2 the chemistry of life Flashcards and Study Sets ...

A B; atom: the basic unit of matter:
nucleus: the center of the atom:
electron: a negatively charged
particle: element: a pure substance
that consists entirely of one type
of atom

Quia - Chapter 2: The Chemistry of Life Vocabulary Review

Chemistry 1405 Chapter 2 1.
chemistry. Chemistry is the study
of matter, its properties, how and
why substances combine or
separate to form other substances,
and how substances interact with

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energy. 2. matter. The term matter refers to anything that occupies space and has mass—in other words, the “stuff” that the universe is made of. 3. three different levels of matter – The three ...

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- Chemistry 1405 Chapter ...

A compound is a substance formed by the chemical combination of two or more elements in definite proportions. (this means that H_2O is water, but H_3O is not). Compounds have different physical and chemical properties from the elements they are made of

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Structure of Atom Class 11 Notes

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Chemistry Chapter 2 • Discovery of Electron—Discharge Tube Experiment In 1879, William Crooks studied the conduction of electricity through gases at low pressure. He performed the experiment in a discharge tube which is a cylindrical hard glass tube about 60 cm in length. It is sealed at both the ends and ...

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What do you see when you look at
this picture? Is it just a mass of
tangled ribbons? Look closely. It 's
actually a complex pattern of three-
dimensional shapes. It represents
the structure of a common
chemical found inside living cells.
The chemical is a protein called
kinase.

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2. What relationship exists between the mass number of an element and isotopes of that element? 3. Explain the difference between ionic and covalent bonds. 4. Compare and contrast adhesion and cohesion. Below is a guide for Chapter 2 Test

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Chapter 2 chemistry of life.

Dehydration synthesis. Hydrolysis.

Subatomic particles that make up
the at.... The atomic number. A
chemical reaction in which two
molecules are bonded together....

A chemical process that splits a
molecule by adding water. Protons,
Electrons and Neutrons. The
number of protons.

chapter 2 chemistry of life
Flashcards and Study Sets ...

This chapter looks at atoms,
bonds, pH and organic molecules.
Good review of chemistry we see
in microbiology.

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Chapter 2 The Chemistry of Life
Reviewing Key Concepts Class
Date Section Review 2-4
Completion On the lines provided,
complete the following sentences.
1. Chemical reactions that energy
often occur spontaneously. 2.
During a chemical reaction,
chemical bonds are 3. Biological
catalysts, or enzymes, act by
lowering the required for a
reaction. 4.

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