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~~$K_c$  Problems Easily | Chemical~~

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~~Constant Chemical Equilibria and~~

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## Chemical Equilibrium Answers

### CHAPTER 18: Chemical

Equilibrium I. a. Write the chemical equilibrium expression for the following equations.

Include the value of K. (a)  $(g) \rightleftharpoons (g)$

$K=0.1$  (b)  $(g) + O_2(g) \rightleftharpoons (g)$   $K = 1.8 \times 10^{-2}$

Does the reaction favor

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products or reactants? (Will there be mostly products or reactants when it reaches 2. a. Compare the rates of forward and reverse reactions when equilibrium has been reached. b.

North St. Paul-Maplewood

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## Oakdale / Overview

Chapter 18: Chemical Equilibrium.

reversible reaction. chemical  
equilibrium. equilibrium constant.  
chemical equilibrium expression.  
a chemical reaction in which the  
products can react to re-form....  
when the rate of its forward

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Chemical reaction equals the  
rate....

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SECTION 18-1 OBJECTIVES Define chemical equilibrium. Explain the nature of the equilibrium constant. Write chemical equilibrium expressions and carry out calculations involving them.

FIGURE 18-1 When heated, mercury(II) oxide decomposes

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Equilibrium Answers  
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into the elements from which it  
was

CHAPTER 18 Chemical Equilibrium  
558 Chapter 18 Chemical  
Equilibrium CHAPTER 18 What  
You'll Learn You will discover that  
many reactions and processes

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reach a state of equilibrium. You will use Le Châtelier's principle to explain how various factors affect chemical equilibria. You will calculate equilibrium concentrations of reactants and products using the equilibrium constant

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Chapter 18: Chemical Equilibrium  
Chemistry: Chapter 18- Chemical  
Equilibrium. Equilibrium. Value of  
K.  $K_{eq} < 1$ .  $K_{eq} > 1$ . when the  
rates of opposing actions are the  
same. represents the reaction  
progression... -all  $K_{eq}$ 's are

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dependent o... reactants of the  
reaction are favored. products of  
the reaction are favored.

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competently as search for them.

In some cases, you likewise accomplish not discover the revelation chapter 18 review chemical equilibrium that you are looking for.

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a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical system (18.2)  
equilibrium position the relative

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concentrations of reactants and products of a reaction that has reached equilibrium; indicates whether the reactants or products are favored in the reversible reaction (18.2)

Chapter 18 Reaction Rates and

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Chemical Equilibrium SECTION 4  
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following questions in the space  
provided. 1. Match the solution  
type on the right to the

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corresponding relationship  
between the ion product and the  
K<sub>sp</sub> for that solution, listed on the  
left. \_\_\_\_\_ The ion product  
exceeds the K<sub>sp</sub>.

CHAPTER 18 REVIEW Chemical  
Equilibrium

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1. In a bottle of unopened cola, the  $\text{CO}_2$  gas dissolved in the liquid is in equilibrium with the  $\text{CO}_2$  gas above the liquid. The dissolved gas reacts with water molecules in the cola to form carbonic acid, which also dissociates into carbon dioxide

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Equilibrium. Which chemical equation(s) best describe this equilibrium system? a.  $\text{CO}_2(\text{g}) \rightleftharpoons \text{CO}_2(\text{l})$  b.

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SHORT ANSWER Answer the  
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provided. 1. Match the solution

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type on the right to the

corresponding relationship

between the ion product and the

$K_{sp}$  for that solution, listed on the

left. \_\_\_\_\_ The ion product

exceeds the  $K_{sp}$ . CHAPTER 18

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