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15: Acid-Base Titration and pH Jeopardy Template. 1. When you  
conduct an acid-base titration, a. the pH of the solution must go up.  
b. the pH of the solution must go down. c. the pH of the solution  
must be 7.0 at the end point. d. the equivalence point must be  
reached., 2.

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## Chapter 15 Acid Base Titration Ph Practice Test

Baylor, Scott / Chapter 15 Acids and Bases Chapter 15 Test Acids  
And Bases Answers when acids or bases act as acids or bases, they  
form their conjugate and the strength is the inverse of their parent.  
and there are certain strong acids and strong bases. LiOH, NaOH,  
KOH - when act as bases (get rid of H<sup>+</sup> ion) then left with Li Page  
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ACIDS AND BASES 453 SECTION 15-1 OBJECTIVES List five  
general properties of aqueous acids and bases. Name common  
binary acids and oxyacids, given their chemical formulas. List five

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acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius' s theory of ionization. Explain the differences

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## CHAPTER 15 Acids and Bases

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Chapter 15: Acids and Bases Acids and Bases Arrhenius

De fi nitions: acids - compounds that produce an increase in  $[H^+]$

when dissolved in water bases - compounds that produce an

increase in  $[OH^-]$  when dissolved in water Lewis De fi nitions:

acids - electron pair acceptors bases - electron pair donors

Br ø nsted-Lowry De fi nitions:

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## Chapter 15: Acids and Bases Acids and Bases

Chapter 15 Acids Bases Review bases C. They can act as acids or

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bases when reacting with each other D. They can act as acids when reacting with HCl molecules 14) 25.0 cm<sup>3</sup> of sulfuric acid solution reacts with 36.2 cm<sup>3</sup> of 0.225 mol dm<sup>-3</sup> sodium hydroxide solution. The concentration of the Chapter 15 Acids Bases Review - e13components.com Chapter 15: Acids and Bases Acids and Bases

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strong base. weak base. an acid that ionizes completely in a solvent.

an acid that releases few hydrogen ions in aqueous solution. a base

that ionizes completely in a solvent. a base that releases few

hydroxide ions in aqueous solution.

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