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Bases Chemistry 102: Chapter 15
Acids and Bases, A Molecular Look

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(University of Jordan) || Part 1

Chapter 15 (Applications of Aqueous
Equilibria) - Part 1 ~~Chemistry 102:~~

~~Chapter 15 Acids and Bases, A~~

~~Molecular Look (University of Jordan)~~

~~|| Part 2 Acid Base Titration Curves,
pH Calculations, Weak~~ /u0026

~~Strong, Equivalence Point, Chemistry~~

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Electrolytes, Acids, Bases Chapter 14

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LAIR CHEM10 - Chapter 15: Acid-Base

Equilibria Acid Ionization Constant

and pK_{a} Chapter 15 Equilibria-

Lewis acids and Bases Chapter 14

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Bases and Salts - Introduction |

Chemistry | Don't Memorise Easy way
to memorize the 7 strong acids and 6
strong bases

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#27Chapter 16 (Spontaneity, Entropy,
and Free Energy) - Part 1 Acids and
Bases 2--How to identify an Acid or

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~~Answer~~ 8.3 Strong and Weak Acids and
Bases Acids Bases and Salts Acid-Base
Equilibria and Buffer Solutions

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Delocalization of Anionic Charge

Chapter 14 (Acids and Bases) - Part 4

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Acid-Base Equilibria Deprotonation in

Weak Acids Chapter 16 Acid-Base

Equilibria Acids, Bases and Salts - 2 |

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Chapter 15 Acids And Bases

Chapter 15: Acids and Bases
Acids and Bases. Arrhenius Definitions: acids-
compounds that produce an increase
in $[H^+]$ when dissolved in water.
bases- compounds that produce an
increase in $[OH^-]$ when dissolved in
water Lewis Definitions: acids-

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Answers
electron pair acceptors.

Chapter 15: Acids and Bases Acids
and Bases

Chapter 15 Acids and Bases. STUDY.
PLAY. Chapter 15 main idea. acids
and bases change the color of

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Answers
compound indicators. ... -Lewis acid-base reaction is the formation of one or more covalent bonds between an electron-pair donor and an electron pair acceptor, although the 3 acid-base reactions differ, many compounds can be categorized as acids ...

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Chapter 15 Acids and Bases
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Chapter 15: Acids and Bases 1. Sour
taste 2. The ability to dissolve many
metals 3. The ability to turn blue
litmus paper red 4. The ability to

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neutralize bases

Chapter 15: Acids and Bases

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ACIDS AND BASES 453 SECTION 15-1

OBJECTIVES List five general
properties of aqueous acids and bases.

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Name common binary acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius' s theory of ionization. Explain the differences

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Acids and Bases. GCh15-1. Chapter 15. Acids and Bases. What we will learn:

- Bronsted acids and bases
- Acid-base properties of water
- pH - a measure of acidity
- Strength of

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Answers

- Weak acids (bases) and acid (base) ionization constant
- Diprotic and polyprotic acids
- Molecular structure and strength of acids
- Acid-base properties of salts
- Acid-base properties of oxides and hydroxides
- Lewis acids and bases.

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Chapter 15. Acids and Bases

An acid that is a compound of hydrogen, oxygen, and a 3rd element (Usually a nonmetal) -If the polyatomic ion is the most common form and ends in "-ate" change it to "-ic". -If the polyatomic ion has one

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Answers less oxygen than the most common form and ends in "-ite" change it to "-ous". Bitter. When in aqueous solutions, bases have this type of taste (Bath soap)

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acids = hydrochloric, nitric, sulfuric,
phosphoric, ethanoic (acetic),
carbonic. bases = potassium
hydroxide, sodium hydroxide, calcium
hydroxide, magnesium hydroxide.
Bronsted-Lowry acids and bases. acid
- molecule or ion that is a proton

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donor (a proton is a hydrogen ion, H^+)
base - molecule or ion that is a proton
acceptor.

Chapter 15 (Acids and Bases)

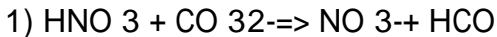
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15 Drill: Identify the B-L acid and base

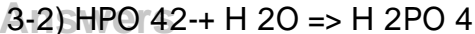
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in each of the following. Circle any amphoteric species acid acid base base Note: In both examples, water behaved as an acid or a base. A species that can act as an acid or a base is called amphoteric. Bronsted-Lowry (B-L) Theory – Cont. acid base



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Bridgewater State University
Acid-Base Properties of Ions and Salts.
-Salts are water-soluble ionic
compounds. -Salts that contain the

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Answers
ation of a strong base and an anion that is the conjugate base of a weak acid are basic. -NaHCO_3 solutions are basic. -Na^+ is the cation of the strong base NaOH .

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Strong acids and bases are often powerful, dangerous substances. (H₂SO₄) will decompose sugar into black charcoal. Nitric acid (HNO₃) reacts with many metals, and hydrochloric acid (HCl) will eat...

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Acids and Bases Chapter 15 Acid
Properties Have a sour taste Acetic
acid, citric acid React with metals to
form H_2 gas React with carbonates
and bicarbonates to form CO_2 gas

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Base Properties Have a bitter taste
Can be slippery, soaps are bases
Ammonia, sodium hydroxide, milk of
magnesia

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Acids and Bases Chapter ...

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The following acid-base reactions go completely in the forward direction: 1. $\text{HCl(aq)} + \text{NH}_3(\text{aq}) \rightarrow \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq})$; (HCl is a strong acid) 2. $\text{HSO}_4^-(\text{aq}) + \text{CN}^-(\text{aq}) \rightarrow \text{HCN(aq)} + \text{SO}_4^{2-}(\text{aq})$; (HSO_4^- is a stronger than HCN) Many acid-base reactions reach a state of equilibrium.

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CHAPTER 15 - ACIDS AND BASES -

Berkeley City College

Chapter 15 - Acids, Bases, & Salts.

Lewis Acid. Lewis Base. Bronsted

Lowry Acid. Bronsted Lowry Base.

Anything that accepts a pair of

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Answers: Anything that donates a pair of electrons in a reaction.

Anything that donates a proton, H^+ , in a reaction. Anything that accepts a proton in a reaction.

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Flashcards and Study Sets ...

Chem 1120 - Chapter 15: Acids and Bases Practice Quiz 1. Match the formulas in Problems #1-4 with the information about them: 1. H_2SO_4 2. CH_3COOH (or HAc) 3. NaOH 4. $(\text{NH}_4)_3\text{PO}_4$. a) used as fertilizer b) found in digestive juices c) present in

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Vinegar d) used in manufacture of soap and paper, and found in Drano

Chem 1120 - Chapter 15: Acids and
Bases - TTU CAE Network

CHAPTER 15: ACIDS AND BASES

Problems: 1-10, 15-18, 51-54, 63-64,

Page 35/40

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67-68, 71-72 15.1 PROPERTIES OF
ACIDS & BASES Acids Bases –
produce hydrogen ions, H^+ –
produce hydroxide ions, OH^- – taste
sour – taste bitter; feel soapy,
slippery – turn blue litmus paper red
– turn red litmus paper blue 15.2
ARRHENIUS ACIDS AND BASES

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CHAPTER 15: ACIDS AND BASES

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Hydronium ion and the

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Chapter 15 Acids and ...

– An acid is a substance that, when dissolved in water, increases the concentration of hydrogen ions. – A

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Answer A base is a substance that, when dissolved in water, increases the concentration of hydroxide ions.

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