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of the quantities of reactants or
products in a chemical reaction
using the relationships found in
the balanced chemical equation.

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sample of a compound having the
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94.07% Sulfur and 5.93%
hydrogen b. 80.68%

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following problems. Assume
excess of any reactant that isn't
mentioned, unless otherwise .
specified. Show your work,
including proper units, to earn full
credit. 1. Given the balanced

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Equation, show what the following molar ratios (i.e., the conversion factors) should be.

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is placed in a 25-mL graduated
cylinder containing 10.0 mL of

water. The level of water rises to
18.0 mL. Aluminum has a density

of 2.7 g/mL. Calculate the mass of
the sample. 2. Saturn is about 1

429 000 km from the Sun.

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Answer Key Answer: 4.93×10
 -5 L or $49.3 \mu\text{L}$ In Example 12.2.1
and Example 12.2.2, the identity
of the limiting reactant has been
apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 ,
ethanol, and para -nitrophenol.
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