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AP Chem - Full
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guide AP Chem

Kinetics Review

Kinetics:

Initial Rates
and Integrated
Rate Laws

AP Chemistry

Unit 5 Part 1

Review: Reaction
Kinetics

Chemical

Kinetics Rate

Laws - Chemistry

Review - Order

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of Reaction

\u0026amp; Equations

AP Chemistry:

**5.1–5.3 Reaction
Rates, Rate Law,
and**

Concentration

Changes AP

~~Chemistry Test~~

~~Prep — Kinetics~~

~~#1 AP Chemistry:~~

~~3.11–3.13~~

~~Spectroscopy,~~

~~Photoelectric~~

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~~Effect, and Beer
Lambert Law AP
Kinetics Notes~~

Differential

Rate Laws *The*

*Cell Cycle (and
cancer)*

[Updated] Le

~~Chatelier's~~

~~Principle of~~

~~Chemical~~

~~Equilibrium~~

~~Basic~~

~~Introduction~~

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Synthesis

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WorkKeys Math

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Questions ~~How to~~

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~~AP Chemistry~~

~~Exam~~ Equilibrium

Equations: Crash

Course Chemistry

#29 ~~Rate Law~~

~~Beers Law~~

~~Reaction Rate~~

~~Laws~~ AP

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Chemistry:

3.4-3.6 Ideal

Gas Law and

Kinetic

Molecular Theory

AP Chemistry Lab

#6 Kinetics of

Hydrogen

Peroxide

Decomposition

Enzymes

(Updated)

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Chemistry's

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Chemistry #32 AP

Chemistry

Investigation

#11: Rate Law of

the Fading of

Crystal Violet.

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AP Chemistry

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Rate Laws For
Reaction
Mechanisms Using
Rate Determining
Step - Chemical
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Response Sample
Questions

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Answer: (a) $A =$
 $abc; 0.600 =$
 $(5000$

$\text{cm}^{-1}\text{M}^{-1}) (1.00$

$\text{cm}) (c) (b) \ln[X]$

$t - \ln[X] 0 =$

$-kt \quad c = 1.20$

$10^{-4} \text{ M} \ln(4.00$

$10^{-5}) - \ln(1.20$

$10^{-4}) = -k(35$

$\text{min}) \quad k = 0.0314$

$\text{min}^{-1} (c) \ln[X]$

$t - \ln[X] 0 =$

$-kt (d) \quad t_{1/2} =$

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$$0.693 k = 0.693 \\ 0.0314 = 22.1 \\ \text{min}$$

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ANSWERS a. For
the first 2
seconds, while

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acceleration is
constant, $d = \frac{1}{2}$
at 1982B1

Substituting the
given values $d =$
10 meters, $t = 2$
seconds gives a
 $= 5 \text{ m/s}^2$ b.

The velocity
after
accelerating
from rest for 2
seconds is given
by $v = at$, so v

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= 10 m/s c.
Response.

Answers

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Response

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ANSWERS

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Answers

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Question 3 5 Br

$2(aq) + BrO_3^-$

$(aq) + 6 H^+(aq)$

$\rightarrow 3 Br_2(l) + 3$

$H_2O(l)$ 3. In a

study of the

kinetics of the

reaction

represented

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above, the
following data
were obtained at
298 K.

Experiment

Essay Questions

1983 - WCS

(b) Calculate
the rate
constant, k , for
the reaction at
350. K. Include
appropriate

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units with your answer. (c) Is the initial rate of the reaction in trial 1 greater than, less than, or equal to the initial rate in trial 2? Justify your answer. (d) The half-life of the reaction in trial 4 is less

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than the half-
life in trial 1.

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QUESTIONS Unit 7
Kinetics and ...
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money for you
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definitely best
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novels, tale,
jokes, and

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Answers: 1. (a)
When we compare
the results of
experiments 2
and 3, we see
that when [B]
doubles, the
rate doubles, so

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the reaction is
first order with
respect to B.

Experiments 2
and 3 prove that
the rate must
double when [B]
doubles. Knowing
this, we can see
that when the
rate doubles
from experiment
1 to 2, it must
be because of B,

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and the

Answers

AP Chemistry

Self-Test

Worksheet

Kinetics

Answer Key

Testname: CH_12_

PRAC_TEST_KINETI

CS.TST MULTIPLE

CHOICE. Choose

the one

alternative that

best completes

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the statement or
answers the
question. 1) C

ID: chem9b

14.1-1 2) B ID:

chem9b 14.1-2 3)

A ID: chem9b

14.1-3 4) E ID:

chem9b 14.1-9 5)

B ID: chem9b

14.1-19 6) B ID:

chem9b 14.1-24

7) E ID: chem9b

14.1-26 8) C ID:

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chem9b 14.1-28

9) D ID: chem9b

14.1-37 10) B

A.P. Chemistry
Practice Test:
Ch. 12, Kinetics

MULTIPLE ...

KINETICS

Practice

Problems and

Solutions d.

Write the rate

law for the

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Response

Answers
reaction. rate =
 $k [A]^2 [B]^2$.

Consider the
following
mechanism. 0 3 ?

$0 \rightarrow 2 + 0$ (fast) 0

$3 + 0 \rightarrow 2 + 0$

(slow) a. Write

the overall

balanced

chemical

equation. $2 \rightarrow 0 + 3$

$? \rightarrow 3 + 0 + 2$ b.

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Identify any intermediates within the mechanism. 0 c.
What is the order with respect to each reactant? 0

KINETICS

Practice

Problems and

Solutions

No credit earned

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for numerical
answer without
justification.

1) a) 5 points .

$$K_p = \frac{(P_{NH_3})^2}{(P_{H_2S})^2 (P_{NH_3})}$$

$$P_{H_2S} = (0.659$$

$$\text{atm} / 2) = 0.330$$

$$\text{atm}. K_p = (0.330$$

$$\text{atm}) (0.330 \text{ atm})$$

$$= 0.109 \text{ atm}^2. \text{ b)}$$

$$5 \text{ points. } P_{NH_3} =$$

$$2 P_{H_2S} (2x) (x)$$

$$= 0.109. x =$$

$$= 0.109. x =$$

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0.233 atm =
 $P_{H_2S} \cdot 2x = 0.466$
atm = P_{NH_3} . c) 5
points.

equilibrium
pressure of NH_3
= equilibrium
pressure of H_2S
= 0.330 atm

Advanced
Placement
Chemistry: 1981
Free Response

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Answers

(a) two points;
one for line of
answer - 232.7
 $J/K = S^{\circ} (C_2H_6)$
- (261.4 +
200.9) $J./K. S^{\circ}$
 $(C_2H_6) = 229.6$
 $J/K. units$
ignored; 1 point
earned for 98.9
 J/K ; 1 point
lost if
stoichiometry is

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not implied in
process (b)
three points
total; one point
each portion;
any value for T
(e.g., 273 K or
298 K) is
allowable:

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curriculum,
simply because
as yet we having
nothing else to
go on, but
having said
that, all of
these ideas are
likely to still
remain relevant
going forward.

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reaction ...

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Rates and
Mechanisms of
Chemical
Reactions 7 For
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our example data
table, the rate
stays the same
regardless of
the
concentration of
[A], therefore,
it is zero order
with respect to
A. However, the
rate doubles
with a doubling
of [B] and
triples with a

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tripling of [B].

Answers

AP* Chemistry

CHEMICAL

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Response

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Practice #4 (p.
49-50) Homework
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Practice

Questions ...

Answer the following questions related to the kinetics of chemical reactions.

$I^{-}(aq) + ClO^{-}(aq) \rightarrow IO^{-}(aq) + Cl^{-}(aq)$ Iodide ion, I^{-} , is oxidized to

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hypoiodite ion,
 IO^- , by
hypochlorite,
 ClO^- , in basic
solution
according to the
equation above.
Three initial-
rate experiments
were conducted;
the results
shown in the
following table.

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