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 $3A(g) + B(g) \rightleftharpoons 2C(g)$ If 2.50 moles of A and 0.500 moles of B are added to a 2.00 L

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container, an equilibrium is established in which the $[C]$ is found to be 0.250 M.

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needed to react with 500.0 g of H_2SO_4 according to the following equation:
 $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$

ANS: 10.19 mol
2. Calculate the mass of NH_3 that can be produced from the reaction of 125 g of NCl_3

Stoichiometry Calculation Practice

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reactions in solution are carried out in the same manner as we discussed in Chapter 11. Instead of masses, however, we use volumes of solutions of known concentration to determine the number of moles of reactants. Whether we are dealing with volumes of solutions of reactants or masses of reactants, the

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Answer Key
coefficients in the balanced chemical equation tell us ...

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12.2 Chemical Calculations 39. Explain the term mole ratio in your own words. When would you use this term? 40.

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